



**FUCHUN SECONDARY SCHOOL  
END OF YEAR EXAMINATION 2017  
SECONDARY 3 EXPRESS**

NAME:

CLASS:

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INDEX  
NUMBER:

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**SCIENCE CHEMISTRY**

**5076, 5078  
Max mark: 85  
5 OCTOBER 2017  
1 h 45 mins**

Additional materials: Multiple Choice Answer Sheet

**READ THESE INSTRUCTIONS FIRST**

Do not use staples, paper clips, highlighters, glue or correction fluid.

**Section A**

Write in soft pencil.

Write your name, class and index number on the Answer Sheet in the spaces provided.

There are **twenty** questions on this section. Answer **all** questions. For each questions there are four possible answers **A, B, C** and **D**.

Choose the one you consider correct and record your choice in **soft pencil** on the OTAS provided.

Each correct answer will score one mark. A mark will nt be deducted for a wrong answer.

Any rough working should be done in this booklet.

**Section B**

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs or rough working.

Answer **all** questions in the spaces provided.

**Section C**

Answer **any two out of three** questions.

Write our answers in the spaces provided on the question paper.

The number of marks is given in brackets [ ]

A copy of the Periodic Table is printed on page 19.

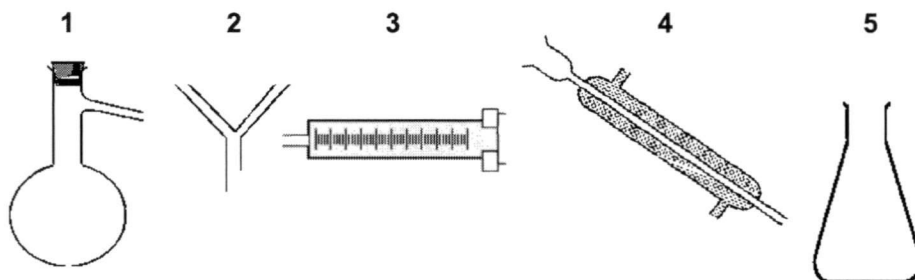
Section	Marks
A	
B	
C	
C	

Name of setter: Mr Nor Mohamad

**This document consists of 19 pages.**

**Section A (20 marks)**  
**Answer all questions.**

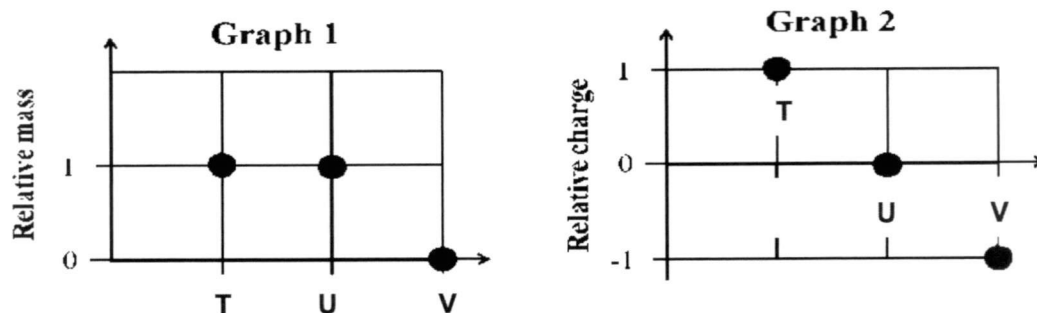
1. The diagram shows some laboratory apparatus.



Which apparatus are needed to produce and collect pure water from seawater?

- A** 2 and 5  
**B** 3 and 5  
**C** 1, 2 and 4  
**D** 1, 4 and 5
2. Ester is a sweet smelling organic substance which is commonly used as artificial flavouring. Which property of an ester could be used to check its purity?
- A** boiling point  
**B** colour  
**C** smell  
**D** solubility in water
3. The relative atomic mass of naturally occurring chlorine is **not** a whole number. What is the reason for this?
- A** Chlorine atoms can have different number of electrons.  
**B** Chlorine atoms can have different number of electron shells.  
**C** Chlorine atoms can have different number of neutrons.  
**D** Chlorine atoms can have different number of protons.

4. Given below are the relative masses and charges of subatomic particles **T**, **U** and **V**.



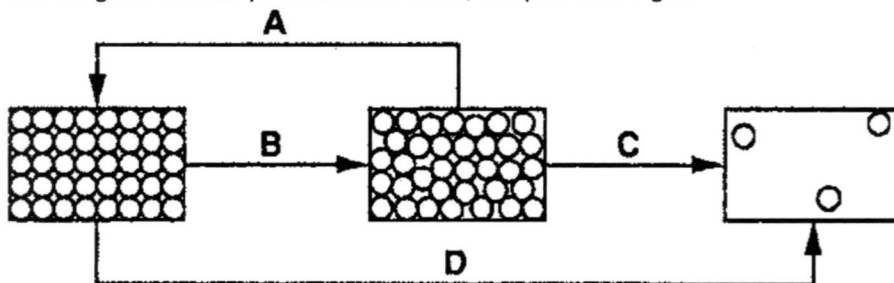
Which of the following correctly describes the identity of particles **T**, **U** and **V**?

	<b>T</b>	<b>U</b>	<b>V</b>
<b>A</b>	proton	neutron	electron
<b>B</b>	electron	proton	neutron
<b>C</b>	proton	electron	neutron
<b>D</b>	electron	neutron	proton

5. The table gives data about four substances.  
In which substance are the particles vibrating in a fixed position at 200 °C?

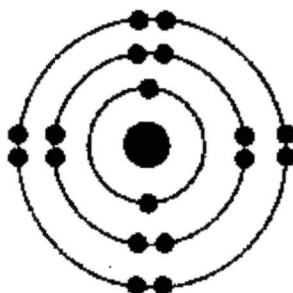
	melting point/ °C	boiling point/ °C
<b>A</b>	- 258	- 183
<b>B</b>	- 124	35
<b>C</b>	121	345
<b>D</b>	253	926

6. The diagram shows particles in a solid, a liquid and a gas.



Which arrow represents sublimation?

7. Which of the following has the electronic structure shown in the diagram below?



- A  $Al^{3+}$   
B  $Cl^{-}$   
C  $Na^{+}$   
D Ne
8. Which **two** statements about an ionic bond are correct?
- 1 It can be formed between two non-metal atoms.
  - 2 It can be formed between a metal atom and a non-metal atom.
  - 3 It can be formed by sharing electrons between the atoms.
  - 4 It can be formed by the transfer of electrons between atoms.
- A 1 and 3  
B 1 and 4  
C 2 and 3  
D 2 and 4
9. Which group of substances contains an element, a mixture and a compound?
- A air, pure water, potassium carbonate  
B air, cobalt, zinc sulfate  
C magnesium, pure water, sulfur  
D sodium chloride, sulfur, zinc
10. A container of distilled water contains dissolved carbon dioxide. Which is the pH of this distilled water?
- A 5  
B 7  
C 9  
D 11

11. The chart shows the colour ranges of four different indicators. Which indicator is blue in an acidic solution?

indicator	pH value													
	1	2	3	4	5	6	7	8	9	10	11	12	13	14
A	— colourless —————><—————								blue —————					
B	— red —————><—————						blue —————							
C	— red —————><—————						blue —————><—————				yellow —————			
D	— yellow —————><—————				blue —————									

12. Which substance is used to increase the pH in a soil?

- A ammonium nitrate
- B calcium hydroxide
- C magnesium sulfate
- D potassium chloride

13. Element L burns in air to form a product that dissolves sparingly in water to give an alkaline solution. What is element L?

- A bromine
- B carbon
- C magnesium
- D neon

14. Which quantity is the same for one mole of argon and one mole of carbon monoxide measured at room temperature and pressure?

- A mass
- B number of atoms
- C number of molecules
- D volume

15. A chemist carried out an experiment to find the relative molecular mass ( $M_r$ ) of a compound. The  $M_r$  of the compound was 28. Which one of the following could be the compound?
- A  $\text{CH}_4$
  - B  $\text{C}_2\text{H}_4$
  - C  $\text{C}_2\text{H}_6$
  - D  $\text{CO}_2$
16. What is the relative molecular mass of hydrated copper (II) nitrate,  $\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$ ?
- A 111
  - B 143
  - C 242
  - D 10152
17. Which of the following is the correct chemical formula of manganese (IV) oxide?
- A  $\text{Mn}_4\text{O}$
  - B  $\text{MnO}_2$
  - C  $\text{Mn}_2\text{O}_4$
  - D  $\text{MnO}_4$
18. Which statement about the elements in Group I of the Periodic Table is correct?
- A The proton (atomic) number is 1 greater than that of the element above it.
  - B They have increasing boiling point down the group.
  - C They become less reactive down the group.
  - D They form chlorides with similar formula.

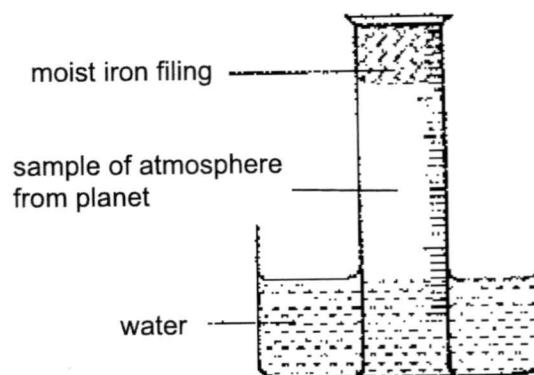
19. Which of the following does **not** correctly describe the trend on descending Group VII?

- A Down Group VII, colour darkens.
- B Down Group VII, melting point increases.
- C Down Group VII, reactivity increases.
- D Down Group VII, density increases.

20. The atmosphere of a newly discovered planet contains the following gases.

carbon dioxide	20%
nitrogen	40%
noble gases	10%
oxygen	30%

The apparatus below was set up with a 100 cm<sup>3</sup> sample of the atmosphere of the planet in the graduated tube. The volume of the sample is measured at intervals until no further change in volume took place.



What volume of the sample will remain?

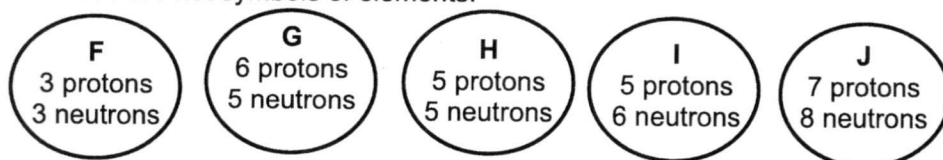
- A 10 %
- B 30%
- C 70%
- D 80 %

**-End Of Section A-**

### Section B (45 marks)

Answer all questions. Write your answers in the spaces provided.

1. The diagram below shows the nuclei of five different atoms. The nuclei labelled **F**, **G**, **H**, **I** and **J**. These are not symbols of elements.



Which letter(s) from **F**, **G**, **H**, **I** and **J** represent(s) [5]

- (a) the nucleus of an atom with an atomic number of six, .....
- (b) the nucleus of an atom with a relative atomic mass of six, .....
- (c) two nuclei from different isotopes of the same element, ..... and .....
- (d) the nucleus of an atom with one electron in its outermost shell, .....
- (e) the nucleus of an atom which forms a negatively charged ion. ....

2. The table shows some information about substance **A** to **E**. [4]

substance	melting point /°C	boiling point /°C	Does it conduct electricity when it is a solid?	Does it conduct electricity when it is a liquid?
<b>A</b>	– 95	69	no	no
<b>B</b>	146	730	no	yes
<b>C</b>	35 – 50	367 – 410	no	no
<b>D</b>	1240	2100	yes	yes
<b>E</b>	1650	2230	no	no

- (a) Which substance is most likely a metal? .....
- (b) Which substance is a liquid at room temperature? .....
- (c) Which substance is a mixture? .....
- (d) Which substance is most likely to have a giant ionic structure? .....



3. Name the pieces of apparatus best used to carry out the following procedures. [4]

- (a) Add 150 cm<sup>3</sup> of liquid to a beaker. ....
- (b) Add 25.0 cm<sup>3</sup> of acid to a flask. ....
- (c) Measure the boiling point of a liquid. ....
- (d) Separate a precipitate from a solution. ....

4. Oxides can be classified into four groups based on their reactions with acids and alkalis.

(a) In Fig 4.1 below, complete the name of the groups of oxides. [2]

(b) Classify the oxides listed below into the correct column. [4]

aluminium oxide      calcium oxide      carbon monoxide      zinc oxide  
copper (II) oxide      sodium oxide      phosphorous (V) oxide      sulfur trioxide

Oxides				
reaction with acids or alkali	reacts only with acids	reacts only with alkalis	reacts with both acids and alkalis	no reaction
type of oxide				neutral
examples of oxides				

Fig 4.1

(c) Give the chemical formula of these oxides: [2]

aluminium oxide : .....

copper (II) oxide : .....

5. Aspirin is a medicine that is used as a painkiller. It is salicylic acid. A student makes a sample of aspirin. He thinks it contains some impurities.

- (a) The student uses chromatography to produce a chromatogram, as shown in Fig 5.1. He uses his own aspirin, pure samples of aspirin and salicylic acid.



Fig. 5.1

Using the chromatogram in Fig. 5.1, comment on the purity of the student's aspirin sample. Explain your answer. [2]

.....

.....

- (b) Salicylic acid is soluble in ethanol.

	melting point / °C	boiling point / °C
salicylic acid	159	211
ethanol	-141	78

Another student wrongly mixes a bottle of ethanol with some salicylic acid and realises her mistake. She decides to use the method shown in Figure 5.3 to separate the two substances.

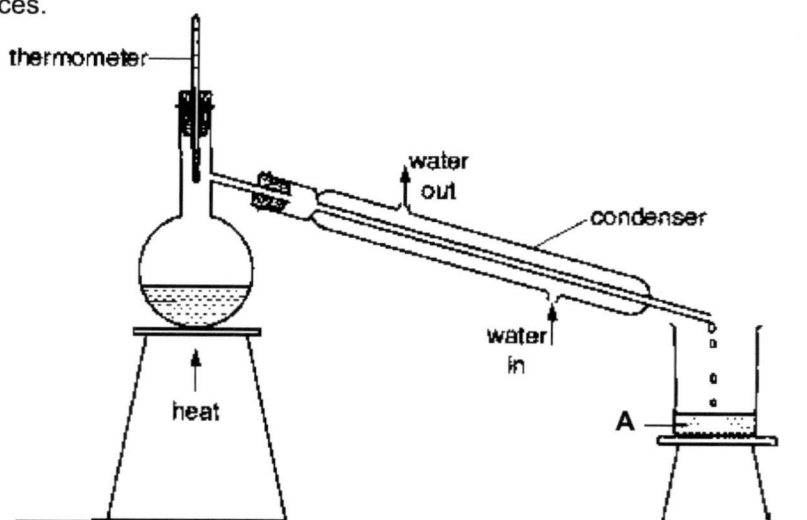


Fig 5.3

- (i) State the general name for this method of separation. [1]

.....

- (ii) Which substance, salicylic acid or ethanol, will be collected at point A?  
Explain your answer [2]

.....

.....

6. Clean air is a mixture of gases

- (a) Name the gas that makes up 0.97% of clean, dry air. [1]

.....

- (b) On cooling, the gases in clean air may liquefy. Describe what happens to the **arrangement** and **movement** of the particles in these gases as they become liquids. [2]

change in arrangement .....

.....

change in movement .....

.....

- (c) Methane and carbon monoxide are two air pollutants. Give a source of each pollutant and describe one of the problems that each can cause. [4]

- (i) methane

source : .....

problem : .....

- (ii) carbon monoxide

source : .....

problem : .....

7. When potassium burns with oxygen, potassium forms potassium ion while oxygen forms oxide ion.

Each potassium ion has a 1+ charge while each oxide ion has a 2– charge.

(a) Draw a 'dot and cross' diagram to show the electronic structure of potassium oxide. [2]

(b) Use your diagram in (a) to explain how and why this change has taken place. [3]

.....

.....

.....

.....

.....

.....

8. Zinc is reacted with dilute nitric acid. Effervescence was observed and the gas produced was collected in the laboratory.

(a) Name the apparatus used to collect and measure the gas produced. [1]

.....

(b) (i) A solution is made by dissolving 12.6 g of nitric acid in water and making the volume up to 500 cm<sup>3</sup>. Calculate the concentration, in g/dm<sup>3</sup>, of this solution. [1]

concentration = .....g/dm<sup>3</sup>

- (ii) A solution contains 12.6 g of nitric acid,  $\text{HNO}_3$  in  $2 \text{ dm}^3$ . Calculate the concentration of this solution in  $\text{mol/dm}^3$ . [1]  
[Relative atomic masses:  $A_r$ : H, 1; N, 14; O, 16]

concentration = ..... $\text{mol/dm}^3$

- (c) (i) Write a balanced chemical equation for the reaction between zinc and nitric acid. State symbols are not required. [2]

.....

- (ii) What is the volume of gas produced when 4.5 g of zinc is reacted with excess dilute nitric acid? [2]  
[The volume of one mole of any gas is  $24 \text{ dm}^3$  at room temperature and pressure]

volume = ..... $\text{dm}^3$

**-End Of Section B-**

**Section C (20 marks)**  
**Answer two out of three questions.**

1. (a) Nitrogen can combine with hydrogen atoms to form ammonia.
- (i) Name the type of chemical bond found in a molecule of ammonia. [1]
- .....
- (ii) Draw a "dot and cross" diagram to show the bonding in a molecule of ammonia. [2]
- .....
- (b) Explain why
- (i) ammonia is a gas at room temperature and pressure, [2]
- .....
- .....
- .....
- .....
- (ii) gaseous ammonia cannot conduct electricity. [1]
- .....
- .....
- (c) Ammonia is a colourless pungent gas. Describe a positive test for ammonia. [1]
- test : .....
- observation: .....
- .....

- (d) Ammonia can be produced when ammonium chloride is reacted with another chemical. Name a chemical that reacts with ammonium chloride to produce ammonia gas, and state the condition required. Write down a balanced chemical equation for the reaction. [3]

chemical : .....

condition required : .....

chemical equation : .....

2. (a) The Periodic Table contains an element with proton number 10 and another element with proton number 18.

- (i) Give the electronic structure of these **two** elements. Use these to explain why both elements appear in the same group of the Periodic Table. [2]

electronic structure of element with proton number 10 : .....

electronic structure of element with proton number 18 : .....

.....

- (ii) Explain why both elements are chemically unreactive. [1]

.....

.....

- (ii) Name one of these two elements, and state one use of this element. [1]

name : .....

use : .....

- (b) Sea water contains potassium bromide, KBr.

Gaseous chlorine can be bubbled into sea water to displace the bromine.

- (i) Describe and explain the observation. [2]

.....

.....

.....

- (ii) Write a balanced chemical equation for the reaction.  
State symbols are **not** required. [1]

.....

- (iii) Suggest another halogen that can be used to displace bromine. [1]

.....

- (iv) 18 dm<sup>3</sup> of chlorine gas is used to displace bromine from sea water.  
Calculate the mass of bromine displaced. [2]

mass of bromine = .....g



3. Magnesium sulfate has been shown to be an effective aid in the fight against blemishes and acne when applied to problematic areas. If combined with water and made into a cream, it can be applied on the face to remove blackheads.

(a) Magnesium sulfate can be produced by reacting magnesium carbonate with a suitable acid.

(i) Identify the acid used. [1]

acid used : .....

(ii) Describe the steps taken to produce a clean and dry magnesium sulfate crystals. [4]

steps : .....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

(b) (i) Write a balanced chemical equation for the reaction in (a).  
State symbols are **not** required. [2]

.....

(ii) Describe a positive test for the gas formed. [1]

test : .....

observation: .....

.....

- (c) Name another substance, other than magnesium carbonate, that can react with the named acid from (a) to prepare magnesium sulfate. [1]

.....

- (d) The method in (a) cannot be used to prepare barium sulfate. Explain why. [1]

.....

**-End of Paper-**

61

57	La	58	Ce	59	Pr	60	Nd	61	Pm	62	Sm	63	Eu	64	Gd	65	Tb	66	Dy	67	Ho	68	Er	69	Tm	70	Yb	71	Lu
lanthanum 139		cerium 140		praseodymium 141		neodymium 144		promethium —		samarium 150		eurogium 152		gadolinium 157		terbium 159		dysprosium 163		holmium 165		erbium 167		thulium 169		ytterbium 173		kisium 175	
89	Ac	90	Th	91	Pa	92	U	93	Np	94	Pu	95	Am	96	Cm	97	Bk	98	Cf	99	Es	100	Fm	101	Md	102	No	103	Lr
actinium —		thorium 232		protactinium 231		uranium 238		neptunium —		plutonium —		americium —		curium —		berkelium —		californium —		einsteinium —		fermium —		mendelevium —		nobelium —		lawrencium —	

## FCS 2017 3E EOY Science Chemistry

### Mark Scheme

#### Section A (20 marks)

Q1	Q2	Q3	Q4	Q5
D	A	C	A	D
Q6	Q7	Q8	Q9	Q10
D	B	D	B	A
Q11	Q12	Q13	Q14	Q15
D	B	C	D	B
Q16	Q17	Q18	Q19	Q20
C	B	D	C	C

#### Sections B and C (Max: 65 marks)

*Notation used in marking:*

- ECF: Error Carry Forward
- BOD: Benefit of Doubt

#### Section B (45 marks)

##### QUESTION 1

- 1 (a) G [1]  
1 (b) F [1]  
1 (c) H and I [1]  
1 (d) F [1]  
1 (e) J or G [1]

##### QUESTION 2

- 2 (a) D [1]  
2 (b) A [1]  
2 (c) C [1]  
2 (d) B [1]

##### QUESTION 3

- 3 (a) measuring cylinder [1]  
3 (b) pipette [1]  
3 (c) thermometer [1]  
3 (d) filter funnel [1]

## QUESTION 4

4

Oxides				
<i>reaction with acids or alkali</i>	<i>reacts only with acids</i>	<i>reacts only with alkalis</i>	<i>reacts with both acids and alkali</i>	[2 m for all correct] [1 m if 2 are correct and 0m if only one is correct]
<i>type of oxide</i>	basic	acidic	amphoteric	
	Calcium oxide Copper (II) oxide Sodium oxide	Phosphorous (V) oxide Sulfur trioxide	Aluminium oxide Zinc oxide	Carbon monoxide
	[1m for every 2 correct oxides classified. No half marks awarded]			

4(iii) aluminium oxide :  $\text{Al}_2\text{O}_3$  [1m]

Copper (II) oxide :  $\text{CuO}$  [1m]

## QUESTION 5

5(a) The student's aspirin is not pure [1m] as there is another point that doesn't belong to the pure aspirin or salicylic acid [1m] {or} it is made of a mixture pure aspirin, salicylic acid and another unknown substance. [1m]

5(b)(i) simple distillation

5(b)(ii) ethanol [1m] , as ethanol has a lower boiling point of  $78^\circ\text{C}$  and will be distilled first [1m]

## QUESTION 6

6(a) Argon [1m]

(b)(i) the particles are in a disorderly (randomly) arrangement; close to each other [1m]

{Many missed out on the disorderly arrangement hence mark not awarded}

(ii) the particles are moving slower. [1m]

(c)(i) decay of vegetation (plant or animal matter)/ waste gases from livestock [1m]

Green house gas / trap energy from the sun which causes global warming. [1m]

(c)(ii) incomplete combustion of fuel [1m]

$\text{CO}$  binds to red blood cell {RBC} which cause the RBC not to be able to transport  $\text{O}_2$  in the body. /lead to breathing difficulties/ respiratory problems [1m]

### QUESTION 7

7(a) . [1m] for 2 correct cations drawn / place a "2" in front of the  $K^+$  ion.

. [1m] for the correct oxide ion drawn.

7(b) K has one electron in the outermost shell which it loses to form the ion. [1m]

O has six electrons in the outemost shell hence it gains 2 electrons to form the ion[1m]

The ions formed has a completely filled valence electron shells made up of 8 valence electrons. [1m] {many missed out on the third point}

### QUESTION 8

8(a) gas syringe . [1m]

8(b)(i) 25.2. [1m]

8(b)(ii) 0.1. [1m]

8(c)(i)  $Zn + 2 HNO_3 \rightarrow Zn(NO_3)_2 + H_2$  . [1m for correct formula] . [1m for balanced eqn]

8(c)(ii) 0.15 mol of Zn. [1m]

0.15mol of  $H_2$  produced

$3.6 dm^3$  of  $H_2$ . [1m]

### SECTION C (Max: 20 Marks)

#### QUESTION 1

1 (a) (i) covalent bond [1m]

(a)(ii) 2m for correct number of electrons shared, and correct number of atoms.

1(b)(i) ammonia is a simple covalent molecule. There is only weak intermolecular forces of attractions [1m] which require low heat to overcome (Not broken). . [1m]

(b)(ii) there is no free ions or free moving electrons to conduct electricity

Hold a damp red litmus paper at the mouth of the test tube.

The damp red litmus paper will turn blue. . [1m]

{many missed out on "damp" hence mark not awarded}

(d) sodium hydroxide or calcium hydroxide . [1m]

Warm/ Heat the mixture. [1m]

$NH_4Cl + NaOH \rightarrow NaCl + NH_3 + H_2O$ . [1m]

## QUESTION 2

2(a)(i) 2.8 ; 2.8.8 [1m]

Both have 8 electrons in the outermost shell or both have 8 valence electrons [1m]

2(a)(ii) The outermost shell is completely filled . [1m]

2(a)(iii) neon –making neon lights

Argon – filling up light bulbs

Helium – filling up balloons / blimp [1m for any correct ]

2(b)(i) the yellowish –green gas will disappear and the solution will turn orange/reddish brown [1m]

Chlorine is more reactive than bromine. [1m]

2(b)(ii)  $2\text{KBr} + \text{Cl}_2 \rightarrow 2\text{KCl} + \text{Br}_2$  [1m]

2b(iii) Fluorine. [1m]

2b (iv) no of mole of  $\text{Cl}_2 = 0.75 \text{ mol}$  [1m]

= no of mol of  $\text{Br}_2$  produced.

Mass of  $\text{Br}_2 = 0.75 \times 160$

= 120 g [1m]

## QUESTION 3

3(a)(i) sulfuric acid . [1m]

(ii) Add **excess** magnesium to sulfuric acid ;

Filter off (excess) magnesium carbonate;

Heat the filtrate until it is saturated;

leave it to cool and crystals to form;

Filter/collect the crystals and Dry the crystals with filter paper

[1m for each correct step max 4 m]

3(b)(i)  $\text{H}_2\text{SO}_4 + \text{MgCO}_3 \rightarrow \text{MgSO}_4 + \text{H}_2\text{O} + \text{CO}_2$  [1m for correct formula] [1m for balanced eqn]

3(b)(ii) Bubble the gas into limewater/calcium hydroxide

white precipitate produced in the lime water [1m]

3(c) magnesium (metal)/magnesium oxide/magnesium hydroxide [1m]

3(d) it is non soluble salt [1m]