

NEW TOWN SECONDARY SCHOOL Mid-Year Examination Secondary 1 Express

NAME	
CLASS	INDEX NUMBER

Lower Sec Science Part 2 (Biology)

8 May 2017 0900-1030 1 hour 30 minutes (For Parts 1 and 2)

READ THESE INSTRUCTIONS FIRST

Write your name, register number and class in the spaces provided above. Write in dark blue or black pen. **DO NOT** use staples, highlighters, glue or correction fluid/tape.

Answer both Part 1 and Part 2.

Section A (10 marks)

Each question consists of four possible answers. Select the **most** appropriate answer and record its alphabet in the space provided on page 6.

Each correct answer will be awarded 1 mark. No marks will be deducted for incorrect answers.

Section B (30 marks)

Answer all questions in the spaces provided.

The use of an approved scientific calculator is expected, where appropriate.

For Examine	r's Use
Section A	
Section B	
Total for	
Part 2	

This document consists of 1.

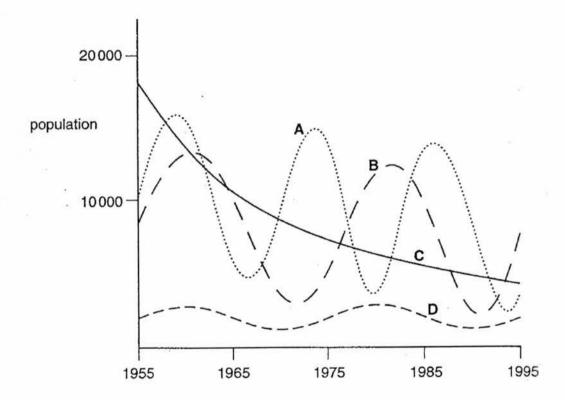
Setter: Mdm S.Valli

Section A (10 marks)

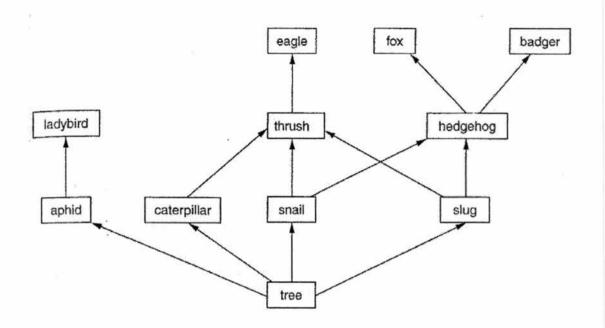
Select the most appropriate answer and record its alphabet in the space provided on page 6.

- Which of the following represents an adaptation of a mangrove plant to oxygen-deficient soil?
 - A They photosynthesise at low tide.
 - B They secrete salt through the leaves.
 - C They have needle like leaves.
 - D They have aerial roots that stick out of the soil.
- 2 The population sizes of four different species of insect were monitored over a period of 40 years. The results are shown on the graph below.

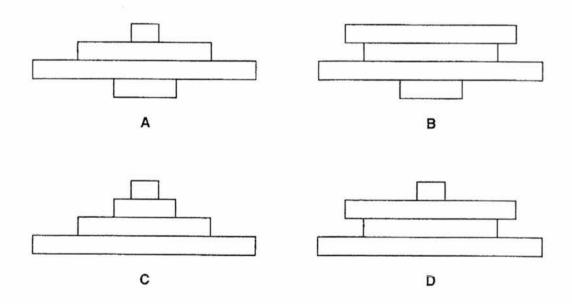
Which species is in the greatest danger of extinction?



3 The diagram shows part of a food web.



Which is a pyramid of energy based on this food web?

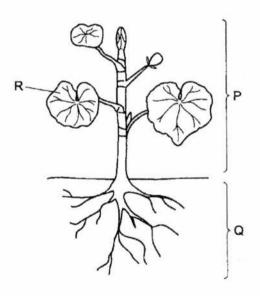


- 4 Humans affect the environment in the following ways.
 - 1. felling of the tropical forest
 - 2. harvesting of marine algae
 - 3. reforestation
 - 4. combustion of fuel
 - 5. overuse of nitrate fertilizer

Which human activities lead to an increase in the level of carbon dioxide in the Earth's atmosphere?

- A 1, 2 and 3
- B 1, 2 and 4
- C 2, 3 and 4
- D 2, 3 and 5

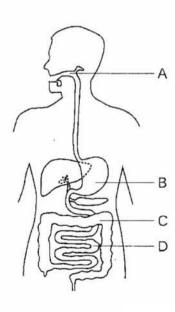
5 The diagram shows a flowering plant



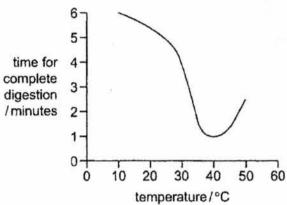
Which correctly identifies P, Q and R?

	Р	Q	R
A	organ	organ	tissue
В	organ	organ system	organ
С	organ system	organ	tissue
D	organ system	organ system	organ

- Which of the following represents the overall magnification of a microscope that has an eyepiece magnification of 10X and an objective magnification of 40X?
 - **A** 10X
 - **B** 40X
 - C 400X
 - **D** 1000X
- 7 Which of the following is true for both xylem vessel and red blood cells?
 - A large surface area to volume ratio
 - B no nucleus
 - C no cytoplasm
 - D thickened cell wall
- 8 Which chemical element forms part of all protein molecules?
 - A calcium
 - B iron
 - C magnesium
 - **D** nitrogen
- 9 The diagram shows the human alimentary canal. In which part do simpler food substances enter the blood stream?



10 The graph shows the effect of temperature on the time taken for the complete digestion of starch.



At which temperature is the rate of digestion of starch the greatest?

- A 10 °C
- B 30 °C
- C 40 °C
- D 50 °C

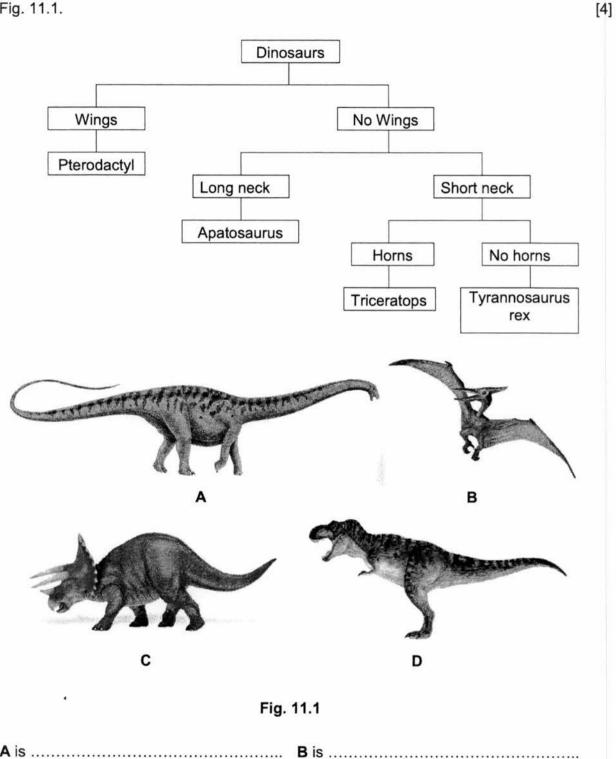
Answers for Section A

1	2	3	4	5	6	7	8	9	10

Section B (30 marks)

Answer all questions in the spaces provided.

11 Identify the different types of dinosaurs shown by using the dichotomous key given in Fig. 11.1.



D is

C is

12 Fig. 12.1 shows the energy flow in kilojoules (kJ) through a food chain.

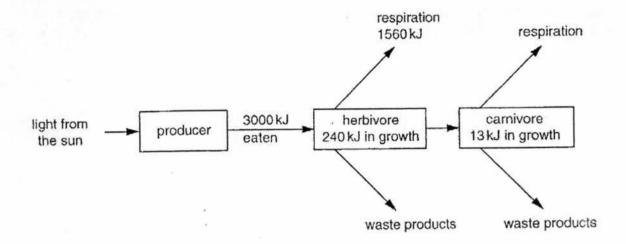


Fig. 8.1

(a)	(i)	How much energy is lost from the food chain as waste products from the herbivores?	[1]
	(ii)	Calculate the percentage of the energy taken in by the herbivore that is used in growth.	[1]
(b)	Sta	ite the importance of chloroplasts in the food chain.	[2]
(c)	Su	ggest why this food chain could not have another trophic level.	[2]

13 Fig. 13.1 shows some muscle cells.

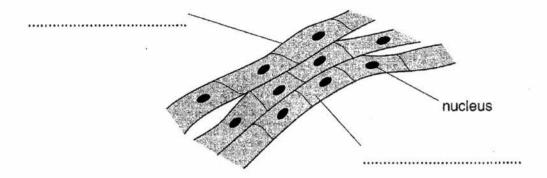


Fig. 13.1

(a)	Complete the labels in Fig. 13.1.	[2]
(b)	Describe two ways Fig. 13.1 shows that these are animal cells and not plant cells.	[2]

(c)	What does a group of muscle cells form?	[1]

14 Some cells are specialised to carry out specific functions.

Fig. 14.1 shows a specialised plant cell.

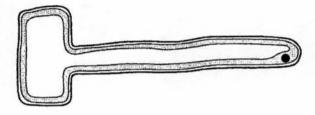


Fig. 14.1

(a)	Name the cell shown in Fig. 14.1.	[1]
(b)	What is the function of the cell shown in Fig. 14.1?	[1]
(c)	Explain how the structure of cell shown in Fig. 14.1 is related to its function.	[2]

15 Fig. 15.1 shows part of the alimentary canal.

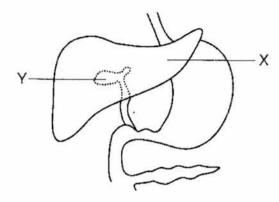


Fig. 15.1

(a)	Name organs A and 1.	[2]
	X	
	Y	
(b)	Describe briefly how organs X and Y increase the efficiency of digestion in the human alimentary canal.	[3]



16

(b)

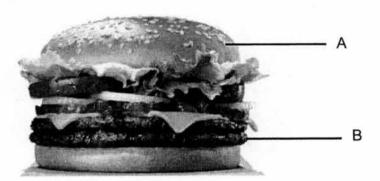


Fig. 16.1

(a) Complete the table below to show the name and nutrient content of the food components of the burger shown in Fig. 16.1. [2]

food component of the burger	put a tick to show the main type of nutrient(s) present
A	carbohydratesproteinsfats
В	carbohydratesproteinsfats

Describe how the food component A and B will be digested in the alimentary canal.	[4]

BP~341

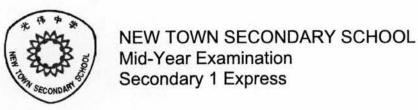
New Town Secondary School 1E Science(Biology) Mid-Year Exam 2017 Marking Scheme

Section A

1	2	3	4	5	6	7	8	9	10
D	C	C	B	D	C	R	D	D	C

Sec	tion B		Marks				
11	B: Pte C: Tric	: Apatosauraus : Pterodactyl : Triceratops : Tyrannasaurus					
12	(a)i	3000 - 1560 - 240 = 1200 kJ (working must be shown)	[1]				
	(a)ii	240/3000 x 100% = 8% (working must be shown)	[1]				
	(b)	the presence of chloroplasts enable green plants to photosynthesise and produce their own food; other organisms feed on plants directly or indirectly and get their energy;	[1] [1]				
	(c)	energy is lost at every trophic level and only a small amount / fraction of energy is passed on to the next trophic level; the amount of energy in the carnivore is too low to be passed on to another organism to sustain it;	[1] [1]				
13	(a)	Cell membrane; Cytoplasm:	[1] [1]				
	(b)	Absence of cell wall; Absence of chloroplasts; Absence of large central vacuole; (any two points)	[1] [1] [1] Max=[2]				
	(c)	muscle tissue	[1]				
14	(a)	root hair cell	[1]				
	(b)	absorb water and mineral salts from the soil	[1]				
	(c)	elongated structure / long protrusion; Increase surface area to volume ratio to increase rate / efficiency of absorption;	[1]				

			BP~342
15	(a)	X: liver Y: gall bladder	[1] [1]
	(b)	liver produces bile; gall bladder stores bile; bile is released into small intestine to emulsify fats which increases the surface area of fats for digestion by enzymes / lipase;	[1] [1]
16	(a)	A: bread / bun Carbohydrates B: meat patty Proteins	[1] [1]
	(b)	carbohydrates partially digested in mouth by amylase into maltose; maltose is completely digested in the small intestine by maltase into glucose; proteins are partially digested in the stomach by protease into	[1] [1]
		polypeptides; polypeptides are completely digested in the small intestine by	[1]
		protease into amino acids;	[1]



NAME		
CLASS	INDEX NUMBER	

Lower Sec Science Part 1 (Chemistry)

8 May 2017 0900-1030 1 hour 30 minutes (For Parts 1 and 2)

READ THESE INSTRUCTIONS FIRST

Write your name, register number and class in the spaces provided above. Write in dark blue or black pen.

DO NOT use staples, highlighters, glue or correction fluid/tape.

Answer both Part 1 and Part 2.

Section A (10 marks)

Each question consists of four options. Choose the best option for each question and write your answer in the **table provided on page 4**.

Each correct answer will be awarded 1 mark. No marks will be deducted for incorrect answers.

Section B (20 marks)

Answer all questions in the spaces provided.

The number of marks is given in [] at the end of each question or part question.

A copy of The Periodic Table is provided on page 9.

The use of an approved scientific calculator is expected, where appropriate.

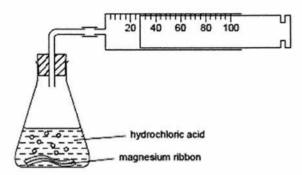
For Examiner	's Use
Section A	
Section B	
Total for	
Part 1	

This document consists of 9

Setter: Mr Mohamad Khirsyaban

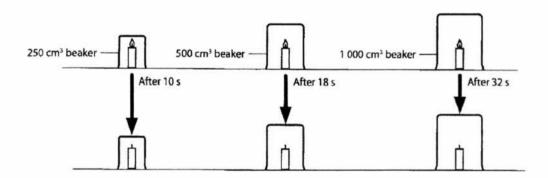
Section A: Multiple Choice Questions (10 marks)

1 A magnesium ribbon is dropped into hydrochloric acid and the gas produced is collected as shown in the diagram below.



Which statement is not an observation made during the experiment?

- A The magnesium ribbon became smaller.
- B Hydrogen gas is produced.
- C Bubbles are produced.
- D 28 cm³ of gas is produced.
- 2 Kelly performed an experiment as shown in the diagram below.



Which hypothesis could Kelly be testing?

- A More oxygen leads to a candle burning longer.
- B The bigger the candle, the longer it burns.
- C A candle stops burning when all the oxygen has been used up.
- D The bigger the beaker, the hotter the temperature of the flame.
- 3 Leonard saw a colourless chemical labelled with the symbol shown below.



What special precaution should he take to when using this chemical?

- A He should heat the liquid before using it.
- B He should wear safety goggles when heating the liquid.
- C He should use a water bath to heat the liquid.
- D He should wear gloves when handling the liquid.

4 Hip joint implants are inserted into patients who have damaged or fractured their hips. These implants, which used to be made of stainless steel, are now manufactured using titanium alloys.

The table below compares the properties of stainless steel and titanium alloys.

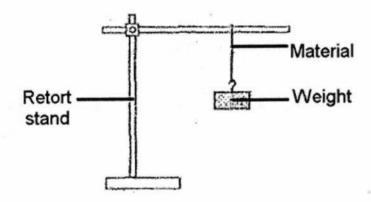
Properties	Stainless steel	Titanium alloy
strength	high	high
density	high	low
resistance to corrosion	high	high
thermal expansion	high	low
cost	low	high
magnetism	low	low

Based on the information, suggest why titanium alloys are used to make hip implants instead of stainless steel.

- A The hip implants are heavier.
- B The hip implants are more expensive to manufacture.
- C The hip implants do not rust or corrode after a long time.
- **D** The hip implants do not change lengths at different temperatures.
- 5 Which of these materials is correctly classified?

	Material	Classification
A	wood	ceramic
В	cotton	fibre
С	diamond	metal
D	steel	plastic

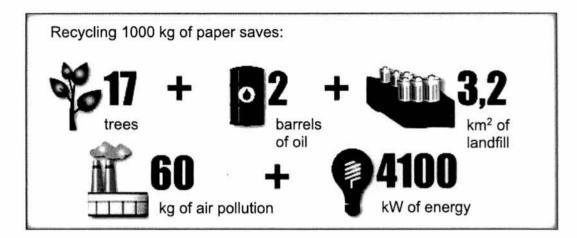
6 Tony set up the apparatus shown below to test the physical property of a material. He recorded the maximum weight he could attach to the material before it breaks.



Which physical property is he investigating?

- A Strength
- B Hardness
- C Flexibility
- D Malleability

7 The diagram below shows some information about recycling paper.



Based on the diagram above, which of the statements below is **not** true about recycling paper?

- A Recycling paper results in fewer forests being cleared.
- B Recycling paper reduces waste.
- C Recycling paper protects the environment.
- D Recycling paper uses up energy.
- 8 Which statement is true about compounds?
 - A They are colourless solutions.
 - **B** They can be separated by physical methods.
 - C They have fixed proportion of elements.
 - D They become a gas at 100 °C.
- 9 Which of the following is **not** true about a suspension?
 - A It is a mixture.
 - B It has the same colour throughout.
 - C It contains both solid and liquid particles.
 - D It is formed when substances cannot dissolve.
- 10 What is the chemical symbol for potassium?
 - A K
 - B Km
 - C P
 - D Po

Answers for Section A

Question	Q1	Q2	Q3	Q4	Q5	Q6	Q7	Q8	Q9	Q10
Answer										

Section B: Structured Questions (20 marks)

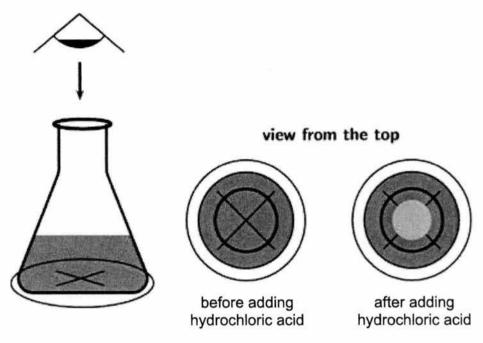
11 The table below shows physical properties of some substances.

substance	appearance	electrical conductivity	melting point
Α	shiny	good	high
В	dull	poor	low
С	transparent	poor	high
D	dull	only conducts when it is a liquid but not solid	high
E	shiny	good	low

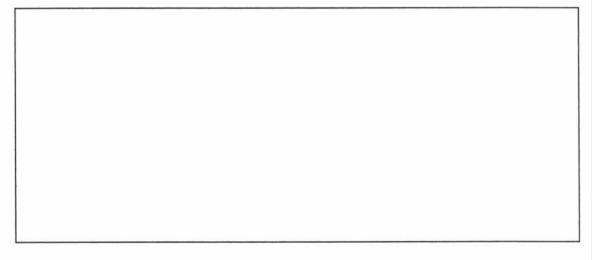
(a)	Which substance in the table above should be used to make wires? Provide reasons for your answer.	
		[2]
(b)	Study the cartoon below carefully.	
	Substance C is plastic. Substance C is plastic. Jillian Jake	
	Suggest a physical property that is missing from the table that might help Jake and Jillian with the identity of substance C . Describe how this property differs for glass and plastic.	
		[2]

Total: [4 marks]

12 Steffi added 25 cm³ of sodium thiosulfate solution into a conical flask, placed on top of a black cross, as shown below. She then added 10 cm³ of hydrochloric acid into the conical flask and found that the cross could not be seen after some time.



(a) Draw and name the apparatus Steffi should use to measure the volume of sodium thiosulfate and hydrochloric acid.



[2]

	ffi wants to investigate whether a higher temperature affects how fast the ss disappears.	
i.	State the independent and dependent variable of this investigation.	
	independent variable:	
	dependent variable:	[2]
ii.	List two variables that she should keep constant for this experiment.	
		[2]
iii.	Should Steffi open or close the air-hole of the Bunsen burner when heating the conical flask? Explain your answer.	
		101
		[2]
	Total: [10 mar	1 7
	Total. [10 mail	rksj
	sh below shows the mass of iodine that can dissolve in 1000 cm ³ of water and	rksj
		rksj
anol a	sh below shows the mass of iodine that can dissolve in 1000 cm³ of water and at different temperatures. Solubility of iodine	rksj
anol a	sh below shows the mass of iodine that can dissolve in 1000 cm³ of water and at different temperatures. Solubility of iodine	rksj
anol a	sh below shows the mass of iodine that can dissolve in 1000 cm³ of water and at different temperatures. Solubility of iodine	rksj
anol a	sh below shows the mass of iodine that can dissolve in 1000 cm³ of water and at different temperatures. Solubility of iodine	rksj
anol a	sh below shows the mass of iodine that can dissolve in 1000 cm³ of water and at different temperatures. Solubility of iodine	rksj
anol a	sh below shows the mass of iodine that can dissolve in 1000 cm³ of water and at different temperatures. Solubility of iodine	rksj
anol a	sh below shows the mass of iodine that can dissolve in 1000 cm³ of water and at different temperatures. Solubility of iodine	rksj
anol a	sh below shows the mass of iodine that can dissolve in 1000 cm³ of water and at different temperatures. Solubility of iodine	rksj
anol a	sh below shows the mass of iodine that can dissolve in 1000 cm³ of water and at different temperatures. Solubility of iodine	rksj
anol a	sh below shows the mass of iodine that can dissolve in 1000 cm³ of water and at different temperatures. Solubility of iodine 700 600 400 300 200 100 0	rksj
anol a	sh below shows the mass of iodine that can dissolve in 1000 cm³ of water and at different temperatures. Solubility of iodine 700 600 400 300 0 100 0 100 200 300 400 500 600	rksj

13

(b)	The	e solute in this experiment is an element.	
	i.	Identify the solute in the experiment above.	
			[1]
	ii.	Explain what is meant by the term element.	
			[1]
(c)		artina spilled some brown iodine solution onto her dress. She soaked her ess in hot water but is unable to remove the stain.	
		ing information from the graph, suggest what she can do to remove the ine stain. Explain your answer.	
			[2]
		Total: [6 ma	arks

The Periodic Table of Elements

							Group	dno								
_	=										=	2	>	N	NII V	0
		2	ž.			-										2
						I										He
			Kev			hydrogen 1										helium 4
	4	protor	proton (atomic) number	number	+						5	9	1	8	6	10
	- G	te	atomic symbol	pol							8	ပ	z	0	4	Ne
Ithium b	peryllium		name								boron 11	carbon	nitrogen	oxygen 16	fluorine	neon
	מפ	reiai	relative atomic mass	MdSS							- 0	7	1 1	2	2 1	40
-	12										13	14	15	16	2	2
e	Ma										A	S	۵.	S	10	Ar
sodium ma	magnesium										aluminium 27	silicon	phosphorus 3.1	Sulfur 32	chlorine	argon
50	20 21	22	23	24	25	26	27	28	29	30	31	32	33	3 8	35	36
, ,	S	F	>	Ö	M	E E	ပိ	Z	ņ	Zn	Ga	Ge	As	Se	B	ž
-	5		vanadium	chromium	mandanese	iron	cobalt	nickel	ropper	zinc	gallium	germanium	arsenic	selenium	bromine	krypton
	40 45	_	51	52	55	99	29	59	64	65	70	73	7.5	79	80	84
37	1	-	41	42	43	44	45	46	47	48	49	20	51	52	53	54
		75	Š	Mo	٦	Ru	쮼	Pd	Ag	В	I	S	Sb	Te	Н	Xe
		*	miobium	molybdenum	technetium	nuthenium	rhodium	palladium	silver	cadmium	indium	ij	antimony	tellurum	iodine	nonex
	88 89		93	96		101	103	106	108	112	115	119	122	128	127	131
1	56 57 - 1		73	7.4	75	76	11	78	79	80	81	82	83	28	85	98
	Ba lanthanoids		Ta	3	Re	SO	11	đ	Au	H	11	Pb	Bi	Po	A	R
caesium	barium	hafnium	tantalum	tungsten	menium	osmium	indium	platinum	gold	mercury	thallium	lead 200	bismuth	mnjuojod	astatine	radon
	137	178	181	184	186	130	192	135	18/	107	204	707	502	1	1	1
13	88 89 - 103	120	105	106	107	108	109	110	111	112		114		116		
ŭ	Ra actinoids	ds	ga	Sg	HB.	£	¥	Ds	Rg	5		FI		^		
francium	radium	Rutherfordium	C	seaborgium	pohríum	hassium	meitnerium	darmstadbum	roentgenium	copernicium	735	flerovium		livermorium		
1	1	1	1	1	1	3	1	1	1	1		1		1		

lanthanoide	22	58	59	09	61	62	63	64	65	99	29	89	69	20	71
Chicipini	Fa	Ce	Ā	PN	Pm	Sm	En	PS	Tb	á	유	ш	Ę	χp	Γn
	lanthanum	cerium	praseodymum	neodynxum	promethium	samarium	europium	gadolinium	terbium	dysprosium	holmium	erbium	thulium	ytterbium	Intetium
	139	140	141	144	1	150	152	157	159	163	165	167	169	173	175
artinoide	89	06	91	92	93	8	95	96	16	98	66	100	101	102	103
collings.	Ac	F	Ра	>	N	Pu	Am	Cm	ă	ಶ	Es	Fm	PW	No.	ב
	actinium	thorium	protactinium	uranium	neptunium	plutonium	americium	conjum	berkelium	californium	einsteinium	fermium	mendelevium	nobelium	lawrencium
	1	232	231	238	1	1	1	1	1	1	1	1	1	ı	1

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

NEW TOWN SECONDARY SCHOOL MID-YEAR EXAMINATION 2017 SECONDARY 1 EXPRESS

BP~353

PART 1: CHEMISTRY ANSWER SCHEME

SECTION A

Question	Q1	Q2	Q3	Q4	Q5	Q6	Q7	Q8	Q9	Q10
Answer	В	А	С	D	В	Α	D	С	В	Α

SECTION B

11(a)	Substance A It is a good electrical conductor and has high melting point.					
	[State both reasons for 1 m]					
11(b)	Flexibility Glass is not flexible while plastic is flexible.	1				
	[Accept other plausible answers]					
12(a)	Measuring cylinder (Drawing of measuring cylinder)	1				
12(b)	Yes her observation is correct. A suspension is formed since the cross cannot be seen after hydrochloric acid has been added. This shows that the mixture is opaque / there are solid particles / particles covering the cross, characteristic of a suspension.	1				
12(c)i	Independent variable: Temperature of mixture Dependent variable: Time taken for cross to disappear					
12(c)ii	Volume of hydrochloric acid / Volume of sodium thiosulfate / Size of cross [Any two variables]	2				
12(c)iii	She should open the air-hole . This will allow her to get a <u>non-luminous flame</u> which is hotter.	1 1				
13(a)	Temperature Nature of solvent	1				
13(b)i	lodine	1				
13(b)ii	A pure substance that cannot be broken down in any simpler substance by chemical means.	1				
13(c)	She should soak the dress in ethanol. Ethanol can dissolve more iodine than water at the same temperature.	1 1				